

Period: \_\_\_\_\_

Name: \_\_\_\_\_

## Nomenclature Practice #1

### Type I Compounds

1. KOH
2.  $\text{NH}_4\text{NO}_3$
3.  $\text{Li}_2\text{O}$
4. NaBr

### Type II Compounds

1. FeO
2.  $\text{Cu}_2\text{O}$
3. HgOH
4.  $\text{PbI}_2$

### Type III Compounds

1.  $\text{CO}_2$
2.  $\text{N}_2\text{O}_5$
3.  $\text{SO}_3$
4. NO

## Nomenclature Practice #2

### Type I Compounds

1.  $\text{MgBr}_2$
2.  $\text{Ca}(\text{OH})_2$
3.  $\text{Na}_2\text{SO}_3$
4. KCl

### Type II Compounds

1.  $\text{FeCO}_3$
2.  $\text{SnBr}_2$
3.  $\text{HgNO}_3$
4. CoSi

### Type III Compounds

1.  $\text{AsCl}_5$
2.  $\text{P}_3\text{N}_5$
3.  $\text{N}_2\text{O}$
4.  $\text{SiI}_4$

## Nomenclature Practice #3

### Acids:

- |                            |                            |                            |
|----------------------------|----------------------------|----------------------------|
| 1. HCl                     | 5. $\text{HClO}_4$         | 9. $\text{H}_2\text{SO}_3$ |
| 2. $\text{HNO}_3$          | 6. $\text{H}_2\text{CO}_3$ | 10. $\text{HClO}_3$        |
| 3. $\text{H}_3\text{PO}_4$ | 7. HF                      | 11. $\text{HNO}_2$         |
| 4. $\text{H}_2\text{SO}_4$ | 8. $\text{H}_3\text{PO}_3$ | 12. HBr                    |

## Nomenclature Practice #4

### Please name the following compounds:

- |                               |                                       |                               |
|-------------------------------|---------------------------------------|-------------------------------|
| 1. $\text{TeI}_4$             | 5. $\text{CdCO}_3$                    | 9. $\text{Mn}(\text{SO}_4)_2$ |
| 2. $\text{Zn}(\text{NO}_2)_2$ | 6. HI                                 | 10. $\text{CoCl}_2$           |
| 3. $\text{Li}_3\text{PO}_4$   | 7. $\text{NiSO}_4$                    | 11. $\text{H}_3\text{PO}_4$   |
| 4. $\text{Al}(\text{OH})_3$   | 8. $\text{AgC}_2\text{H}_3\text{O}_2$ | 12. $\text{K}_2\text{SO}_3$   |